

Process for Removal of SO₂ from Off-Gases by Reaction with H₂O₂

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The present invention relates to a process for removal of SO₂ from off-gases by reaction with H₂O₂.

5 It has been known for more than 30 years that SO₂, as described in the publication 2164e from Lurgi/Südchemie AG, August 1989, can be removed from off-gases by contacting the off-gas in an absorption tower with circulating solution of dilute sulphuric acid containing H₂O₂, whereby SO₂ is dissolved and oxidised to H₂SO₄ in the solution. The
10 circulating solution typically contains 30-60% H₂SO₄ and 0.1-0.5% H₂O₂. The absorption is typically carried out at a temperature of 50-80°C of the circulating solution. H₂O₂ is added either as a concentrated aqueous solution of H₂O₂ to the circulating acid, or it is produced by electrolysis of
15 a side stream of the circulating acid. The produced acid is drawn off from the circulating acid.

The known process usually requires installation of a low velocity aerosol filter downstream of the absorption tower
20 to remove sulphuric acid aerosol (acid mist) in order to meet acid mist emission regulations requiring less than about 5 mol ppm H₂SO₄ in the stack gas. Fine acid mist (aerosol) that may be present in the off-gas is not removed efficiently in the absorption tower. Fine acid mist is also
25 formed in the absorption tower itself by reaction between SO₂ and H₂O₂ vaporised from the absorbing liquid.

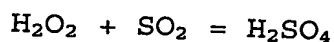
It is a disadvantage of the known process that it requires installation of both an absorption tower and a low velocity
30 mist filter.

In the process according to the present invention, SO_2 in off-gases is removed by reaction with H_2O_2 without the use of an absorption tower by spraying a solution of H_2O_2 in water or dilute sulphuric acid into the off-gas upstream of a low velocity aerosol filter or wet electrostatic precipitator (WESP).

A preferred embodiment of the invention is shown in Fig. 1. A solution of 0.1-30% H_2O_2 in line 1 is sprayed by the spray nozzles 3 into a stream of off-gas in line 2 containing typically between 100-1000 ppm SO_2 and having a temperature typically in the range of 50-120°C. The nozzles are placed in duct 4, so that the spray is evenly distributed in the gas stream upstream of the mist filter 5 in which the gas is passed in parallel through a number of low velocity filter candles 6. Even distribution of the droplets in the gas is desirable for the process and the most even distribution of the droplets is achieved by using air-atomising nozzles producing very small droplets. The H_2SO_4 formed in the process accumulates in the filter elements or candles from which it is drained off through line 7. Most or all of the mass of the droplets evaporate before the gas enters the filter candles, whereby most of the H_2O_2 evaporates and reacts in the gas phase under formation of sulphuric acid aerosol. However, it is not necessary that the droplets are completely evaporated before the gas enters the filter elements. The reaction between SO_2 and H_2O_2 will be completed and the thermal equilibrium will be established in the mist filter elements without decreasing the efficiency of the SO_2 -removal.

Thus, the injection of the aqueous solution H_2O_2 serves two purposes:

5 Firstly, it adds to the off-gas the amount of H_2O_2 , which is required for achieving the desired conversion of SO_2 into H_2SO_4 by the reaction



10 Most of the conversion takes place by reaction in the gas phase between SO_2 and vaporised H_2O_2 under formation of acid mist or between SO_2 and H_2O_2 dissolved in the droplets. The reaction is completed in the aerosol filter in which re-
15 remaining SO_2 is absorbed and reacts with remaining H_2O_2 contained in the dilute sulphuric acid wetting the fibre material.

Secondly, the water comprised in the solution cools off the off-gas in line 2 by evaporation of the droplets, whereby
20 the off-gas is cooled off to a desired temperature of the filter elements or candles typically to a temperature between 50°C and 70°C . The concentration of H_2SO_4 in the produced acid will be the equilibrium concentration of H_2SO_4 at the actual temperature and H_2O partial pressure in the
25 gas phase.

Up to 98% SO_2 -removal can be achieved at, typically, about 95% utilisation of the H_2O_2 .

EXAMPLE

An off-gas stream of 1000 Nm³/h at 100°C contains 200 ppm SO₂ + 10% H₂O and has a temperature of 100°C. 96% SO₂ removal is desired. The aerosol filter is designed for operation at maximum 70°C. Operation at 67-70°C is chosen in order to achieve the highest possible acid strength and low content of remaining H₂O₂ in the produced acid.

The process is conducted as follows: 15.7 kg/h water containing 2.0 wt% H₂O₂ is injected into the off-gas, whereby the off-gas is cooled to 65-70°C in thermal equilibrium. The mist filter is 75 mm thick and has a flow area of 2.5 m². The diameter of the fibres is about 8 µm. Experiments carried out under these conditions show that about 96% of the SO₂ is removed under production of 1.7 kg/h 50% H₂SO₄ with about 500 ppm H₂O₂. The treated gas contains less than 2 ppm H₂SO₄ and the content of H₂O₂ is below the detection limit.

CLAIMS

1. A process for removal of SO_2 in off-gases having a temperature of 30-150°C and containing 0.001-1 vol% SO_2 in which the SO_2 is oxidised to H_2SO_4 by spraying an aqueous solution of H_2O_2 into the off-gas upstream of an aerosol filter removing the produced sulphuric acid from the off-gas.
2. A process as in claim 1, in which the off-gas is cooled by evaporation of the water comprised in the solution being sprayed into the off-gas upstream of the filter.
3. A process as in claim 1, in which a wet electrostatic separator is used in place of an aerosol filter.

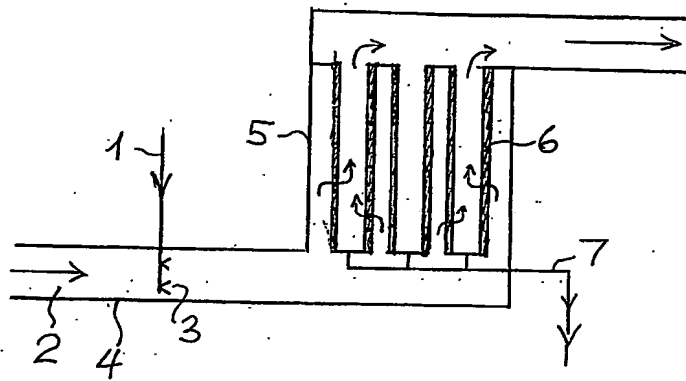


Fig 1

(12) INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

(19) World Intellectual Property
Organization
International Bureau



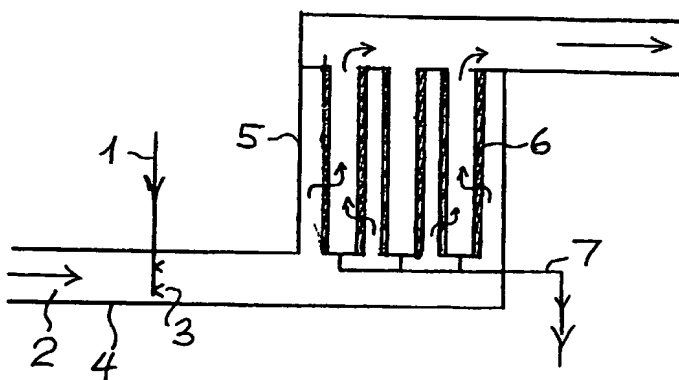
(43) International Publication Date
8 July 2004 (08.07.2004)

PCT

(10) International Publication Number
WO 2004/056449 A3

- (51) International Patent Classification⁷: **B01D 53/50**
- (21) International Application Number:
PCT/EP2003/013699
- (22) International Filing Date: 4 December 2003 (04.12.2003)
- (25) Filing Language: English
- (26) Publication Language: English
- (30) Priority Data:
PA 2002 01992 21 December 2002 (21.12.2002) DK
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- (81) Designated States (*national*): AE, AG, AL, AM, AT, AU, AZ, BA, BB, BG, BR, BW, BY, BZ, CA, CH, CN, CO, CR, CU, CZ, DE, DK, DM, DZ, EC, EE, EG, ES, FI, GB, GD, GE, GH, GM, HR, HU, ID, IL, IN, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MA, MD, MG, MK, MN, MW, MX, MZ, NI, NO, NZ, OM, PG, PH, PL, PT, RO, RU, SC, SD, SE, SG, SK, SL, SY, TJ, TM, TN, TR, TT, TZ, UA, UG, US, UZ, VC, VN, YU, ZA, ZM, ZW.
- (84) Designated States (*regional*): ARIPO patent (BW, GH, GM, KE, LS, MW, MZ, SD, SL, SZ, TZ, UG, ZM, ZW), Eurasian patent (AM, AZ, BY, KG, KZ, MD, RU, TJ, TM), European patent (AT, BE, BG, CH, CY, CZ, DE, DK, EE, ES, FI, FR, GB, GR, HU, IE, IT, LU, MC, NL, PT, RO, SE, SI, SK, TR), OAPI patent (BF, BJ, CF, CG, CI, CM, GA, GN, GQ, GW, ML, MR, NE, SN, TD, TG).
- Declaration under Rule 4.17:**
— of inventorship (Rule 4.17(iv)) for US only
- Published:**
— with international search report
- (88) Date of publication of the international search report:
30 September 2004
- For two-letter codes and other abbreviations, refer to the "Guidance Notes on Codes and Abbreviations" appearing at the beginning of each regular issue of the PCT Gazette.

(54) Title: PROCESS FOR REMOVAL OF SO₂ FROM OFF-GASES BY REACTION WITH H₂O₂



(57) Abstract: A process for removal of SO₂ in off-gases having a temperature of 30-150°C and containing 0.001-1 vol% SO₂ in which the SO₂ is oxidised to H₂SO₄ by spraying an aqueous solution of H₂O₂ into the off-gas upstream of an aerosol filter removing the produced sulphuric acid from the off-gas.

WO 2004/056449 A3

INTERNATIONAL SEARCH REPORT

International Application No
PCT/EP 03/13699

A. CLASSIFICATION OF SUBJECT MATTER

IPC 7 B01D53/50

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

IPC 7 B01D

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the International search (name of data base and, where practical, search terms used)

EPO-Internal, WPI Data

C. DOCUMENTS CONSIDERED TO BE RELEVANT

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A	DE 44 24 367 A (SIEMENS AG) 18 January 1996 (1996-01-18) the whole document	1-3
A	US 3 733 393 A (COUILLAUD J ET AL) 15 May 1973 (1973-05-15) the whole document	1-3
A	DE 40 36 899 A (DEGUSSA) 21 May 1992 (1992-05-21) the whole document	1-3

☐ Further documents are listed in the continuation of box C.

☒ Patent family members are listed in annex.

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Date of the actual completion of the International search

6 April 2004

Date of mailing of the International search report

15.07.2004

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INTERNATIONAL SEARCH REPORT

Internat

Application No

PCT/EP 03/13699

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